

Vanadium Oxide/Graphene Nanoplatelet as a Cathode Material for Mg-Ion Battery

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Abstract

The aim of the present work is to introduce a high performance cathode for magnesium-ion batteries. A simple ball mill process is employed to synthesize $(V_2O_5)_{1-x}$ (Graphene Nanoplatelets (GNP))_x nanocomposite, (where $x = 0, 5, 10, 15, 20$ and 25 wt.% GNP). The synthesized samples are characterized using scanning electron microscope (SEM), X-ray diffraction (XRD) technique, impedance spectroscopy, cyclic voltammetry and charge-discharge test. The maximum conductivity of the investigated samples was found to be 6×10^{-1} S/cm for optimum composite film (25 wt% GNP) at room temperature. Room temperature rechargeable magnesium batteries are constructed from Mg as anode material, $(V_2O_5)_{1-x}$ (GNP)_x as a cathode material and the simple non-aqueous electrolyte based $MgNO_3 \cdot 6H_2O$. Mg/ V_2O_5 cells employing as-prepared electrolyte exhibit initial discharge capacity ~ 100 mAhg⁻¹ while Mg/(V_2O_5 /GNP ($x = 25$ wt.%)) cathode produces a lower initial capacity of ~ 90 mAhg⁻¹. The high initial discharge capacity of V_2O_5 can be attributed to the presence of a large (001) interlayer spacing (~ 11.53 Å) for facile Mg⁺ insertion/extraction.

Keywords

Magnesium Batteries, Vanadium Oxide, Graphene, Conductivity

1. Introduction

Costly Chevrel phase Mo_6S_8 cathode, and complicated electrolyte based Mg [ALCL₂BuET]₂/tetrahydrofuran (volatile) are traditional materials in early rechargeable magnesium batteries, although, these batteries suffer from low voltage, low energy density, kinetic sluggish of Mg⁺² insertion/extraction [1] [2] [3] [4]. Various cathode materials for magnesium battery have been suggested in recent years, such as TiS₂ nanotubes [5], MoS₂ [6], GeO₂ [7], TiS₃ [1], V₂O₅ [8] and MnO₂ [9]. Vanadium pentoxide (V_2O_5)

is a semiconductor material ($E_g = 2.4$ eV), where tuning of fermi level due to ion insertion is expected. The intercalation reaction of Mg in V_2O_5 can be written as follows: ${}_xMg + V_2O_5 \leftrightarrow Mg_x + V_2O_5^-$. During this reaction the vanadium atoms are partially reduced from a V^{5+} to a V^{+4} formal oxidation state [10]. V_2O_5 is belonging to the layered transition metal oxides which possess the ability to structural deformations during the insertion of Mg^{+2} or other bivalent ions. The superior electrochemical performances of V_2O_5 could be ascribed to the unique structure revealing the presence of a large (001) crystal planes interlayer spacing (~ 11.53 Å), which provide large interlayer spacing for facile ion insertion/extraction [11] [12] [13]. V_2O_5 has been studied very intensively as a cathode material for Li-ion batteries [14]. Because the ionic radii of Li^+ and Mg^{2+} are comparable in magnitude, 68 and 65 pm, respectively, the replacement of Li^+ ions by Mg^{2+} ions in insertion compounds is possible. On the other hand, although the electrochemical performance of V_2O_5 has shown great improvement, it suffers from its poor electronic conductivity, which may lead to both of poor capacity and the cyclic ability of V_2O_5 electrodes. Graphene nanoplatelets represent a new class of carbon nanoparticles with multifunctional properties. Graphene nanoplatelet addition can provide barrier properties, while their pure graphitic composition makes them excellent electrical [15] and thermal conductors and can prevent the vanadium dissolution, and alleviate the aggregation of the particles. Since graphene, the name given to a flat monolayer of carbon atoms tightly packed into a two-dimensional (2D) honeycomb lattice, exhibits superior electrical conductivities, high surface areas and chemical tolerance intrigue, many researchers have studied the V_2O_5 /graphene cathode [16] using different preparation methods to control particle size and particle shape, aiming to improve the efficiency of rechargeable batteries. In the present work we will introduce a new rechargeable magnesium battery from electrolyte system based on reaction products of $MgNO_3 \cdot 6H_2O$, succinonitril, tetraethylene glycol dimethyl ether solvent, and $((V_2O_5)_{0.75}/GNP_{0.25})$ nanocomposite cathode. One of our important goals is to reduce the cost of the battery. So the starting materials will be from market and the ball mill process is the desirable one.

2. Experimental

Graphene Nanoplatelets Grade M GNP was characterized by average (7 nm thickness, 10 nm particle diameter, 10^7 S/m Electrical conductivity and surface area $\sim 120 - 150$ m^2/g) have been imported from XG Science company. Composites of V_2O_5 and GNP were prepared by ball milling process under 4 hours' time duration. The resulted product is designated here as $(V_2O_5)_{1-x}(GNP)_x$ composites. The morphology of the nanocomposite was examined using SEM (JOEL-JSM Model 5600). The XRD patterns of the films were taken using Rigaku diffractometer type RINT-Ultima IV/S. The diffraction system based with Cu tube anode with voltage 40 KV and current 40 mA. The current-voltage characteristics of the cathodes were carried out by means of a computer controlled 2400 Keithley electrometer. For electrochemical performance testing, working electrodes (V_2O_5 and V_2O_5 -graphene composite powder) were prepared by mixing

85 wt.% sample as the active material, 6 wt.% conductive agent (carbon black, Super-P-Li), and 9 wt.% poly-vinylidene difluoride (PVDF) binder, N-methylpyrrolidone (Alfa) was then added to produce a viscous slurry and the resultant slurry was pasted onto copper foil. The as-prepared working electrodes were then dried in a vacuum oven at 373 K for 2 h. Electrochemical cells (CR2032 coin type) were assembled in room temperature and ambient pressure by using the working electrode, a separator (filter paper), Mg ribbon as the reference and counter electrode, and 4 gm $\text{MgNO}_3 \cdot 6\text{H}_2\text{O}$ in a 2:10 (w:v) mixture of succinonitril, tetraethylene glycol dimethyl ether, respectively, as the electrolyte. **Figure 1** shows the structure scheme of $\text{MgNO}_3 \cdot 6\text{H}_2\text{O}$, succinonitril and tetraethylene glycol dimethyl ether. Where $\text{MgNO}_3 \cdot 6\text{H}_2\text{O}$ is the Mg^{2+} pump, tetraethylene glycol dimethyl ether is a solvent and succinonitril is a plasticizer agent to dissociate ions and improve ionic conductivity. Cyclic voltammograms (CVs) were conducted in three-electrode cell using an electrochemical instrument of CHI604E Electrochemical Workstation. The cells were charged and discharged on a multi-channel battery test system (NEWARE BTS-TC35) over the voltage range of 0 - 1.6 V versus Mg/Mg^{2+} at constant constant current density $\sim 40 \mu\text{Acm}^{-1}$.

3. Results and Discussion

Figures 2(a)-(c) shows SEM images for V_2O_5 and $\text{V}_2\text{O}_5/\text{GNP}$ nanocomposites. The graphene nanoplatlets are crumpled to a curly and wavy shape. By adding GNP nanoplatlets to V_2O_5 , the later was found to be uniformly distributed in GNP and well distributed on the 2D graphene nanoplatelets, as shown in **Figures 2(a)-(c)**. Moreover, the graphene sheets can prevent the aggregation of V_2O_5 particles to a certain extent, which can be of great benefit to electrochemical reactions. The XRD patterns of the synthesized $(\text{V}_2\text{O}_5)_{1-x}(\text{GNP})_x$ composites (where $x = 0, 5, 10, 15, 20$ and 25 %wt GNP) are shown in **Figure 3**. The XRD spectrum for $x = 0\%$ indicates that the V_2O_5 composites

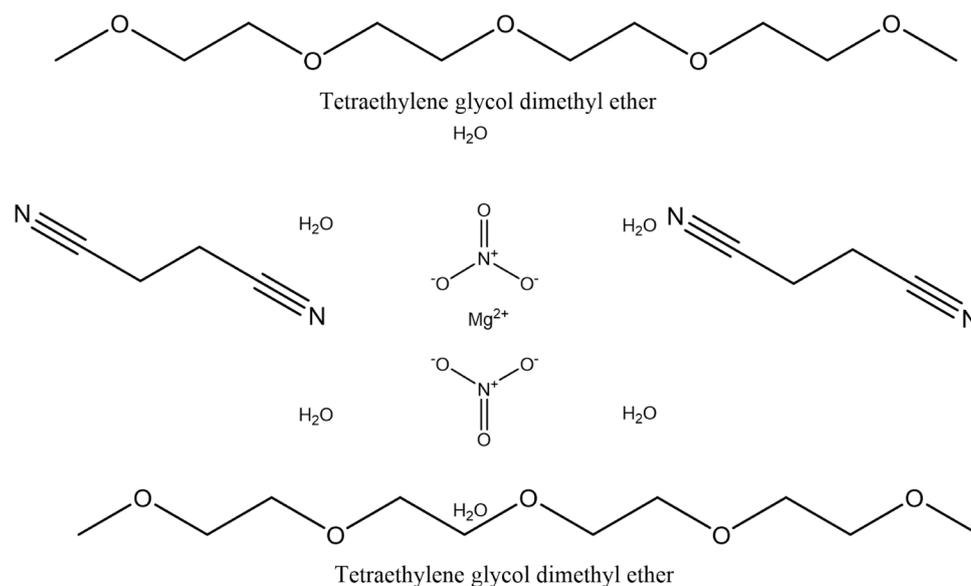


Figure 1. Schematic illustrates the electrolyte structure.

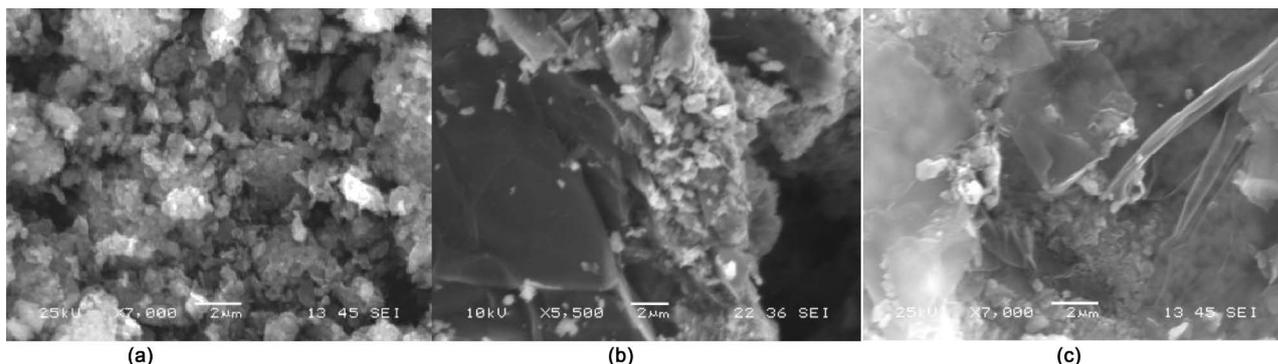


Figure 2. SEM micrographs of $(V_2O_5)_{1-x}(GNP)_x$ composites: (a) 0 wt% GNP; (b) 15 wt% GNP; (c) 25 wt% GNP.

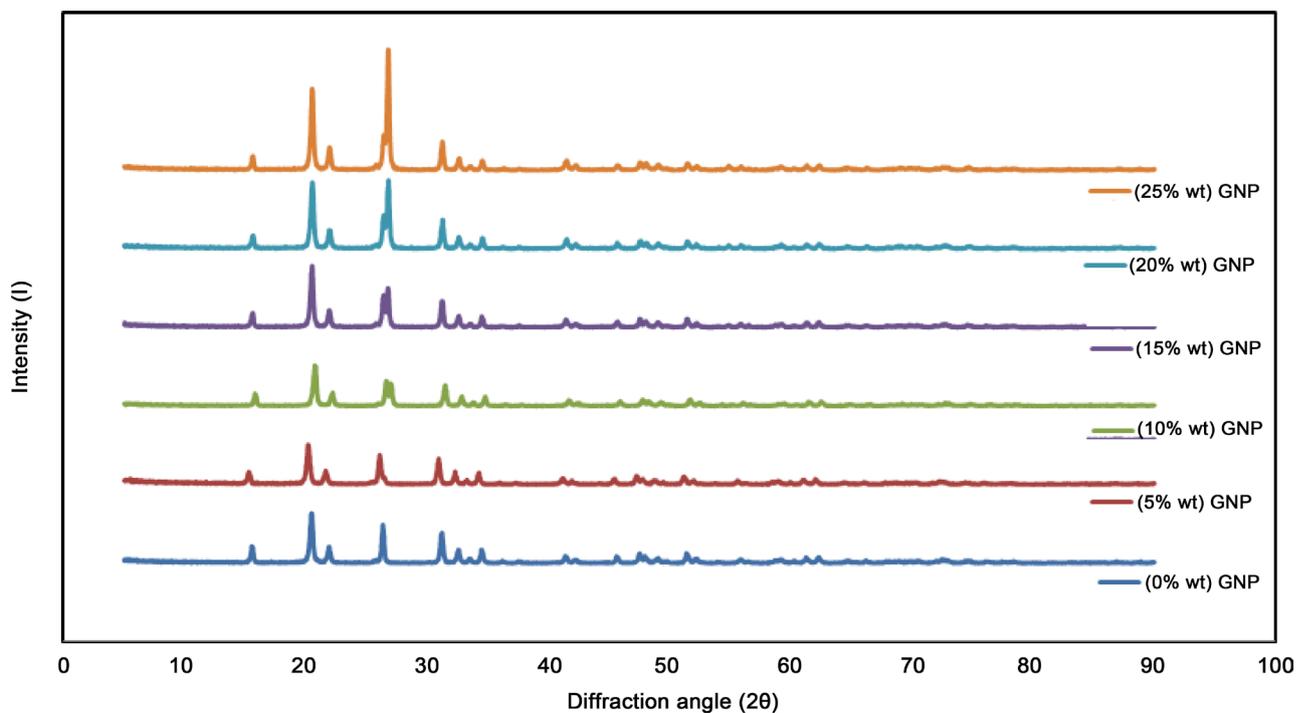


Figure 3. XRD pattern of $(V_2O_5)_{1-x}(GNP)_x$ composites.

are highly crystallized in structure and the entire diffraction peaks match well with Bragg reflections of the pure orthorhombic phase of V_2O_5 nanoparticles, which is consistent with the standard JCPDS No.41-1426 (space group Pmmn) [17] [18]. However, for $x > 0\%$ an additional diffraction shoulder peak around 26.77° , partially overlapping with the V_2O_5 (110) peak ($2\theta = 26.39^\circ$), originates from the (002) diffraction of the graphite. The intensity of the later peak increases linearly when x increases from 5 to 25% and its position corresponds to ~ 31 nm (according to the relation: $\lambda/(2 \cdot \sin\theta)$, $\theta = 26.39^\circ$) spacing between atomic planes [19]. Also, the XRD spectra show a weak peak at 54.8° which corresponds to the (004) reflection of the graphite. It can be noticed that the graphite (002) peak position shifts to the lower angles for $x = 5\%$ and to higher angles for $x = 10\%$. This can be related to the amount of oxygen functional groups formed between the platelets of the graphite. From the XRD measurements, no additional

phases related to structural defects can be detected, while orthorhombic V_2O_5 can be detected after ball milling of V_2O_5 and GNP. Small shift in peak positions of V_2O_5 after doping by GNP was observed which confirm change in (001) interlayer spacing. The crystallite size (τ) of the investigated samples $(V_2O_5)_{1-x}(GNP)_x$ nano-composites (where $x = 0, 5, 10, 15, 20$ and 25 %wt GNP) can be calculated using the first sphere approximation of Debye-Scherrer formula [20] $\tau = K \lambda / \beta \cos \theta$, where K is the shape factor, λ is the X-ray wavelength, β is the line broadening at half the maximum intensity in radians, and θ is the Bragg angle. Using this formula, crystallite dimensions of about 30 nm could be calculated from the high intense peaks. **Figure 4(a)**, shows the current voltage (I&V) characteristics of $(V_2O_5)_{1-x}(GNP)_x$ nano-composite. Generally, the current increases linearly with increasing voltage obeying Ohm law. The dc conductivity was calculated using the equation: $\sigma_{dc} = \frac{1}{R} \times \frac{L}{A}$ [21], where t is the thickness of the sample and A is the surface area of the sample. The value of the resistivity R was measured from the slopes of the straight lines in ohmic region I α V **Figure 4(a)**. The effect of graphene content on the dc conductivity σ_{dc} of $(V_2O_5)_{1-x}(GNP)_x$ is shown in **Figure 4(b)**. The dc conductivity of the composites exhibits insulator behavior for pure V_2O_5 recording $\sim 10 \times 10^{-6}$ S/m, whereas semiconductor behavior for low graphene content up to $x = 10$ wt% was recorded. The electrical conductivity increased as the content of graphene was close to percolation threshold up to $x = 5$ wt%. Above the percolation threshold σ_{dc} was found to increases exponentially before it reach to the saturated point ($x = 10$ wt%) which may be attributed to the formation of filler network and the composites may reach to the metallic behavior. This behavior can be described according to the percolation theory as the following power relation [22],

$$\sigma_{dc} = \sigma_c (V_{GNP} - V_C)^t \text{ for } V_{GNP} > V_C$$

where σ_c is the conductivity of conducting component, V_{GNP} is the volume fraction of Graphene, V_C is the critical volume fraction or percolation threshold, and the exponent t reflects the dimensionality of the system and has been calculated to be either 1.3 or 2.0 corresponding to two or three dimensions, respectively [23]. **Figure 4(b)** (inset) shows a fitting of percolation equation for $(V_2O_5)_{1-x}(GNP)_x$ composites. The exponent t was found to be about 1.4. This result confirms that the GNP nanoparticles are not located on the surface of the host material matrix particles, but it coordinated in the V_2O_5 crystal structure and the formation of graphene three dimensional network will enhancement.

Cyclic voltammetry and discharge curve

Figure 5(a) shows the schematic of the cell configuration of the $Mg^- V_2O_5$ cells in this study. The activity of V_2O_5 and V_2O_5 -graphene nanocomposite for hosting Mg^{+2} ions was evaluated using cyclic voltammetry (CV) and galvanostatic discharge-charge techniques. **Figure 5(b)**, shows the CV results obtained from 2.5 to 0 V using V_2O_5 and V_2O_5 -graphene nanocomposite as a working electrode in a three-electrode cell employing magnesium metal as the counter and reference electrode at a scan rate of 0.05 mV s^{-1} . Although GNP succeeded to increase conductivity of V_2O_5 , it failed to increase

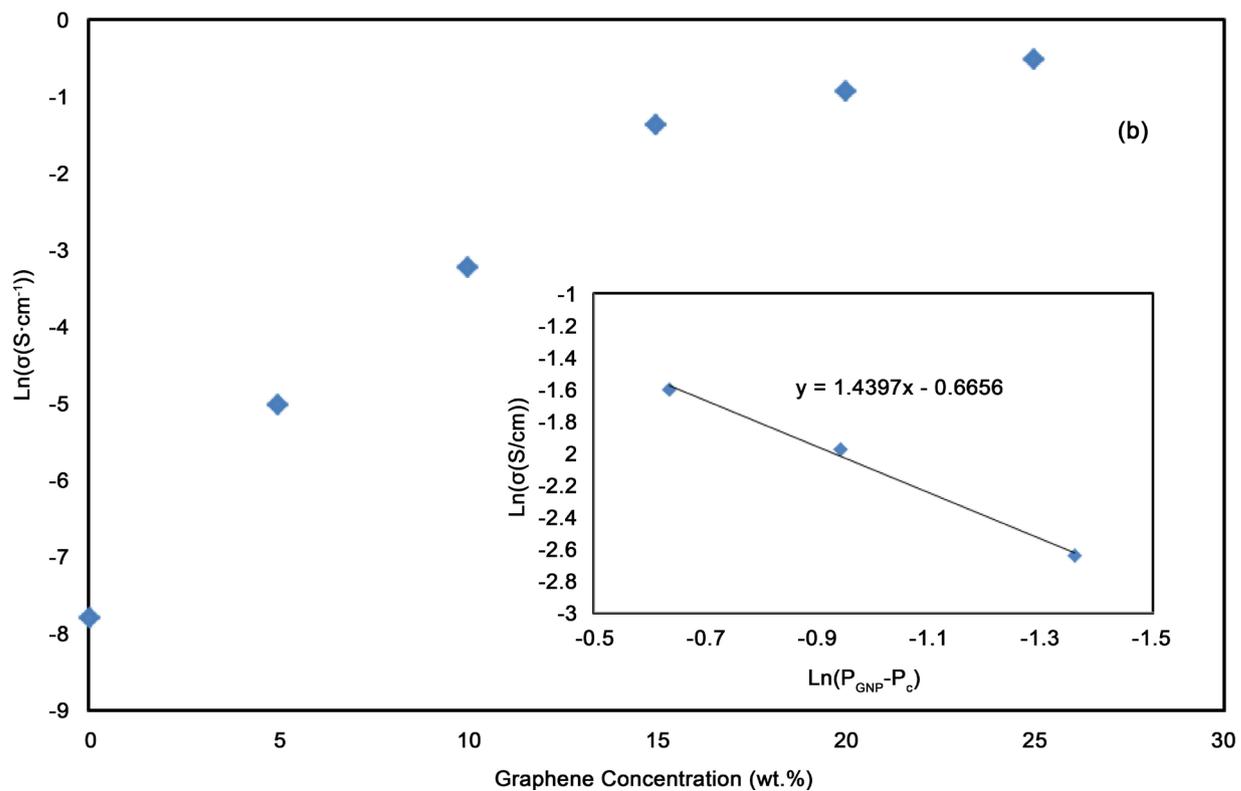
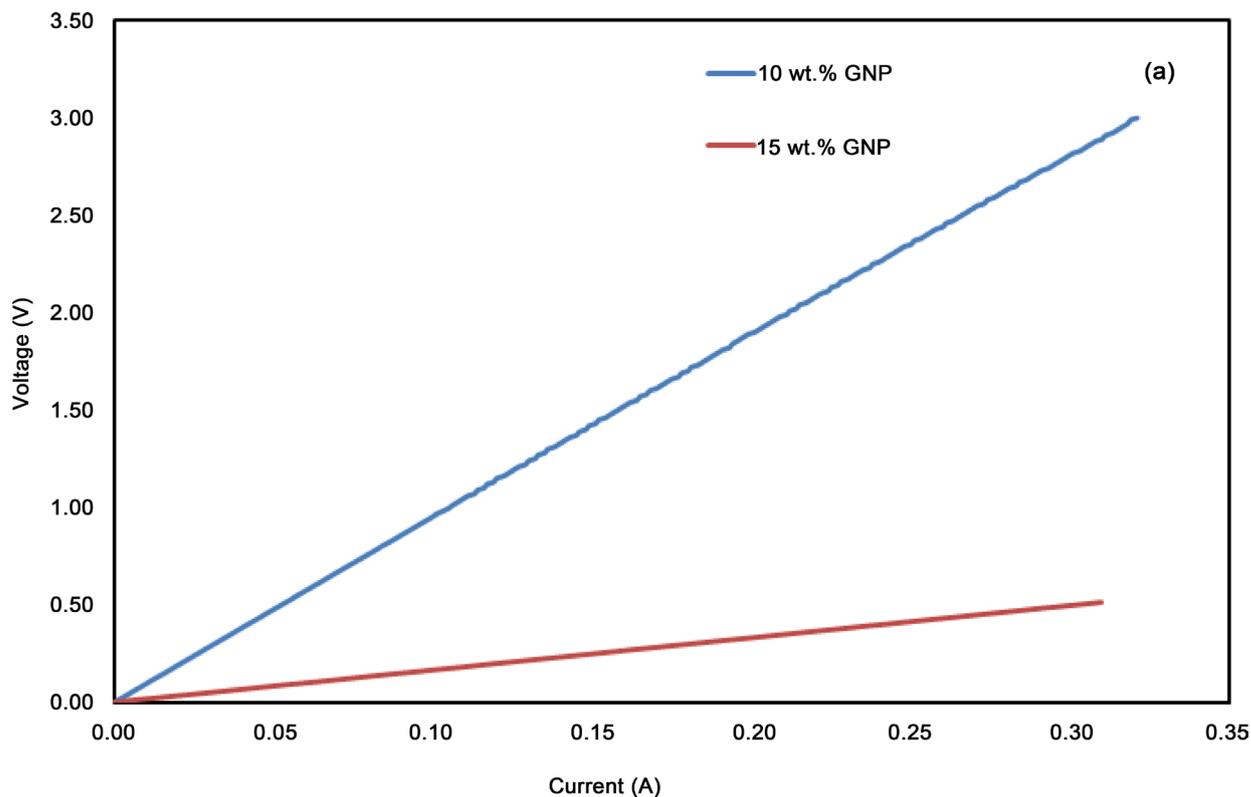


Figure 4. (a) I-V curves of $(V_2O_5)_x/(GNP)_{1-x}$ composites; (b) Variation of dc conductivity of $(V_2O_5)_x/(GNP)_{1-x}$ composites at different graphene concentrations

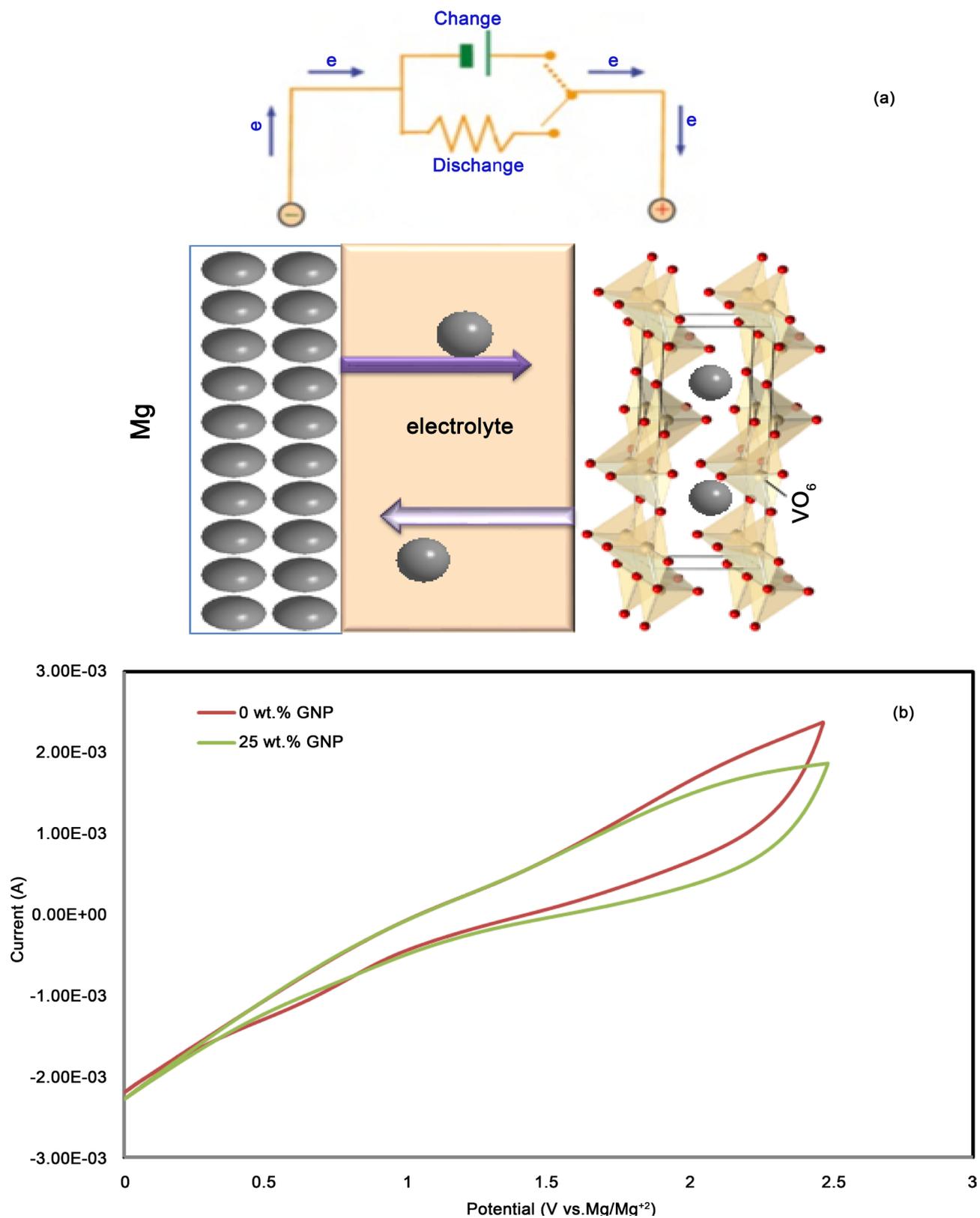


Figure 5. (a) Schematic illustrates Mg^{+2} insertion/extraction within V_2O_5 ; (b) Comparison of the CV curves (at $5\text{ mV}\cdot\text{s}^{-1}$) for V_2O_5 and V_2O_5 /graphene nano-composite.

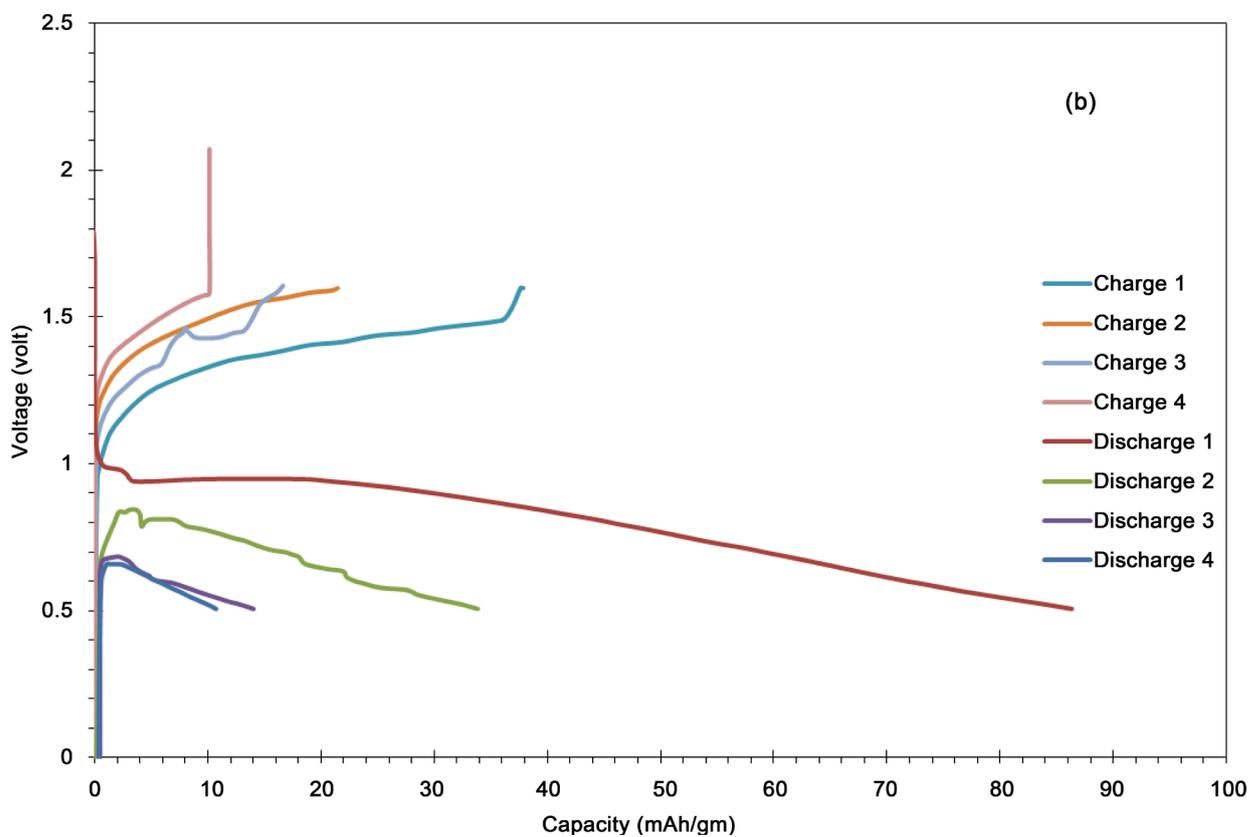
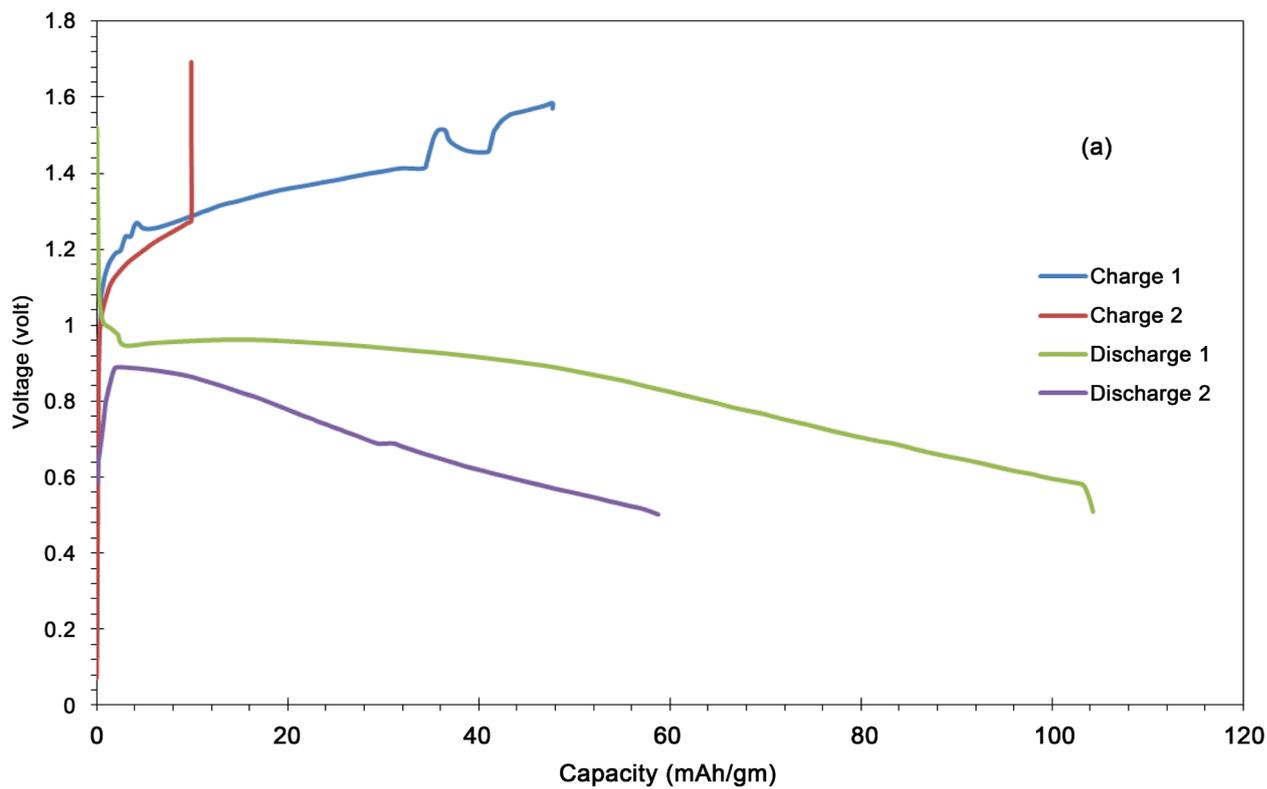


Figure 6. Discharge-charge profiles of (a) V_2O_5 (b) V_2O_5 /graphene nano-composite cathodic materials.

the cathodic and anodic current density. This can be attributed to that the introduction of GNP in V_2O_5 using ball mill technique can perturb the (001) interlayer spacing of V_2O_5 and may be reduce intercalation rate. The perturbation in the interlayer spacing probably decreases the probability of Mg^{+2} insertion/extraction between interlayer spacing and hence decrease the specific capacity as we will see.

Figure 6(a) & **Figure 6(b)** show discharge/charge profiles of Mg/V_2O_5 and $Mg/(V_2O_5/GNP)$ coin cells, in which current density was fixed at $=40 \mu Acm^{-1}$ and the cells were discharged to 0 V and charged to 1.6 V. The initial discharge capacity of Mg/V_2O_5 and $Mg/(V_2O_5/GNP)$ coin cells are approximately 100 and 90 $mAhg^{-1}$, respectively. The combination of V_2O_5 with GNP decrease the discharge capacity compared to pure V_2O_5 . We could not obtain more than 2 and 4 cycles for pure and graphitized V_2O_5 , respectively. We think the structure property of V_2O_5 which give the advantage (001 large interlayer spacing ($\sim 11.53 \text{ \AA}$)) of facile Mg^+ insertion/extraction was deformed after initial cycling. The intercalation of Mg^{+2} may perturb the bonding scheme of V_2O_5 and losses it this property.

4. Conclusion

In summary, V_2O_5/GNP cathode was synthesized by a ball mill method. The integration of V_2O_5 and graphene nanoparticles enhanced the electrical performances. This improved performance could be attributed to the formation of framework nanoscale electrode of 2D graphene decorated with well-dispersed V_2O_5 nanoparticle. Although GNP perturbed the (001) interlayer spacing hence, it failed to enhance the electrochemical performance. As-prepared V_2O_5 and V_2O_5/GNP cathodes can deliver a high capacity of 100 and 90 $mAh \cdot g^{-1}$ respectively, which provides a new direction to explore cathode materials for rechargeable Mg batteries. Further extensive investigations are required, however, to raise the performance of the magnesium-based rechargeable cells to practical levels.

References

- [1] Kouji, T., Yunpeng, G., Yukari, K., Takafumi, Y. and Hidenori, T. (2016) Rechargeable Mg Battery Cathode TiS_3 with d-p Orbital Hybridized Electronic Structures. *Applied Physics Express*, **9**, Article ID: 011801.
- [2] Su, S.J., Huang, Z.G., NuLi, Y., Tuerxun, F., Yang, J. and Wang, J.L. (2015) A Novel Rechargeable Battery with a Magnesium Anode, a Titanium Dioxide Cathode, and a Magnesium Borohydride/Tetraglyme Electrolyte. *Chemical Communications*, **51**, 2641-2644. <http://dx.doi.org/10.1039/C4CC08774G>
- [3] Aurbach, D., Lu, Z., Schechter, A., Gofer, Y., Gizbar, H., Turgeman, R., Cohen, Y., Moshkovich, M. and Levi, E. (2000) Prototype Systems for Rechargeable Magnesium Batteries. *Nature*, **107**, 724-727. <http://dx.doi.org/10.1038/35037553>
- [4] Zhang, M., MacRae, A.C., Liu, H. and Meng, Y.S. (2016) Communication—Investigation of Anatase- TiO_2 as an Efficient Electrode Material for Magnesium-Ion Batteries. *Journal of the Electrochemical Society*, **163**, A2368-A2370. <http://dx.doi.org/10.1149/2.1091610jes>
- [5] Tao, Z., Xu, L., Gou, X., Chen, J. and Yuan, H. (2004) TiS_2 Nanotubes as the Cathode Materials of Mg-Ion Batteries. *Chemical Communications*, **18**, 2080-2081.

- <http://dx.doi.org/10.1039/b403855j>
- [6] Hsu, C.-J., Chou, C.-Y., Yang, C.-H., Lee, T.-C. and Chang, J.-K. (2015) MoS₂/Graphene Cathodes for Reversibly Storing Mg²⁺ and Mg²⁺/Li⁺ in Rechargeable Magnesium-Anode Batteries. *Chemical Communications*, **52**, 1701-1704. <http://dx.doi.org/10.1039/C5CC09407K>
- [7] Sheha, E. and Bassyouni, A. (2016) Structure, Thermal and Electrical Properties of Germanium Oxide/Graphene Nano-Composite Cathode for Magnesium. *Energy and Environment Focus*, **5**, 1-6.
- [8] Niya, Sa., Kinnibrugh, T.L., Wang, H., Gautam, G.S., Chapman, K.W., et al. (2016) Structural Evolution of Reversible Mg Insertion into a Bilayer Structure of V₂O₅·nH₂O Xerogel Material. *Chemistry of Materials*, **28**, 2962-2969. <http://dx.doi.org/10.1021/acs.chemmater.6b00026>
- [9] Lu, W., Karina, A., Per, E.V., Ann, M.S. and Fride, V.-B. (2016) Sponge-Like Porous Manganese (II, III) Oxide as a Highly Efficient Cathode Material for Rechargeable Magnesium Ion Batteries. *Chemistry of Materials*, **28**, 6459-6470. <http://dx.doi.org/10.1021/acs.chemmater.6b01016>
- [10] Laubach, S., Schmidt, P., Thipen, A., Fernandez-Madrigal, F., Wu, Q. and Jaegermann, W. (2007) Theoretical and Experimental Determination of the Electronic Structure of V₂O₅, Reduced V₂O_{5-x} and Sodium Intercalated NaV₂O₅. *Chemical Physics*, **9**, 2564-2576.
- [11] Su, D. and Wang, G. (2013) Single-Crystalline Bilayered V₂O₅ Nanobelts for High-Capacity Sodium-Ion Batteries. *ACS Nano*, **7**, 11218-11226. <http://dx.doi.org/10.1021/nn405014d>
- [12] Gershinshy, G., Yoo, H.D., Gofer, Y. and Aurbach, D. (2013) Electrochemical and Spectroscopic Analysis of Mg²⁺ Intercalation into Thin Film Electrodes of Layered Oxides: V₂O₅ and MoO₃. *Langmuir*, **29**, 10964-10972. <http://dx.doi.org/10.1021/la402391f>
- [13] Chiku, M., Takeda, H., Matsumura, S., Higuchi, E. and Inoue, H. (2015) Amorphous Vanadium Oxide/Carbon Composite Positive Electrode for Rechargeable Aluminum Battery. *ACS Applied Materials & Interfaces*, **7**, 24385-24389. <http://dx.doi.org/10.1021/acsami.5b06420>
- [14] Yu, R., Zhang, C., Meng, Q., Chen, Z., Liu, H. and Guo, Z. (2013) Facile Synthesis of Hierarchical Networks Composed of Highly Interconnected V₂O₅ Nanosheets Assembled on Carbon Nanotubes and Their Superior Lithium Storage Properties. *ACS Applied Materials & Interfaces*, **5**, 12394-12399. <http://dx.doi.org/10.1021/am4033444>
- [15] Liu, H. and Yang, W. (2011) Ultralong Single Crystalline V₂O₅ Nanowire/Graphene Composite Fabricated by a Facile Green Approach and Its Lithium Storage Behavior. *Energy & Environmental Science*, **4**, 4000-4008. <http://dx.doi.org/10.1039/c1ee01353j>
- [16] Du, X., Huang, G., Qin, Y. and Wang, L. (2015) Solvothermal Synthesis of GO/V₂O₅ Composites as a Cathode Material for Rechargeable Magnesium Batteries. *RSC Advances*, **5**, 76352-76355. <http://dx.doi.org/10.1039/C5RA15284D>
- [17] Venkatesan, A., KrishnaChandar, N., Arjunan, S., Marimuthu, K.N., MohanKumar, R. and Jayavel, R. (2013) Structural, Morphological and Optical Properties of Highly Monodispersed PEG Capped V₂O₅ Nanoparticles Synthesized through a Non-Aqueous Route. *Materials Letters*, **91**, 228-231. <http://dx.doi.org/10.1016/j.matlet.2012.09.117>
- [18] Chen, D., Yi, R., Chen, S., Xu, T., Gordin, M.L., Lv, D. and Wang, D. (2014) Solvothermal Synthesis of V₂O₅/Graphene Nanocomposites for High Performance Lithium Ion Batteries. *Materials Science and Engineering: B*, **185**, 7-12. <http://dx.doi.org/10.1016/j.mseb.2014.01.015>
- [19] Karteri, I., Karatas, S. and Yakuphanoglu, F. (2014) Electrical Characterization of Graphene

- Oxide and Organic Dielectric Layers Based on Thin Film Transistor. *Applied Surface Science*, **318**, 74-78. <http://dx.doi.org/10.1016/j.apsusc.2014.01.013>
- [20] Chand, N., Rai, N., Agrawal, S.L. and Patel, S.K. (2011) Morphology, Thermal, Electrical and Electrochemical Stability of Nano Aluminium-Oxide-Filled Polyvinyl Alcohol Composite gel Electrolyte. *Bulletin of Materials Science*, **34**, 1297-1304. <http://dx.doi.org/10.1007/s12034-011-0318-7>
- [21] Gamal, R., Sheha, E., Shash, N. and El-Shaarawy, M.G. (2015) Effect of Tetraethylene Glycol Dimethyl Ether on Electrical, Structural and Thermal Properties of PVA-Based Polymer Electrolyte for Magnesium Battery. *Acta Physica Polonica A*. **127**, 803-810. <http://dx.doi.org/10.12693/APhysPolA.127.803>
- [22] Fan, Y., Wang, L., Li, J., Li, J., Sun, S., Chen, F., Chen, L. and Jiang, W. (2010) Preparation and Electrical Properties of Graphene Nanosheet/Al₂O₃ Composites. *Carbon*, **48**, 1743-1749. <http://dx.doi.org/10.1016/j.carbon.2010.01.017>
- [23] Pang, H., Chen, T., Zhang, G., Zeng, B. and Li, Z.M. (2010) An Electrically Conducting Polymer/Graphene Composite with a Very Low Percolation Threshold. *Materials Letters*, **64**, 2226-2229. <http://dx.doi.org/10.1016/j.matlet.2010.07.001>



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